

KENYAPLEX EXAMS
END TERM 3 2019
FORM 3
CHEMISTRY PAPER 3
CONFIDENTIAL INSTRUCTIONS.

In addition to the apparatus and fittings found in a chemistry laboratory each candidate will require the following:

- Solid P (1.8g of oxalic acid) in a stoppered container.
- Solution Q – 100cm³
- Pipette and pipette filler.
- 100 cm³ measuring cylinder.
- 250cm³ beakers (two)
- 250cm³ volumetric flasks.
- 1 label
- Burette
- One boiling tube
- Five dry clean test tubes in a rack
- Clamp stand
- Solution L – 50cm³
- Stop watch
- 10cm³ measuring cylinder (2)
- Tripod stand and wire gauze
- Test tube holder
- Funnel
- 250 cm³ conical flasks (2)
- Test tube holders
- Distilled water
- White tile
- Solid D – 2g Lead II Nitrate in a stoppered container.
- Solid M – 2g Sodium Chloride in a stoppered container.
- Solid T – 2g Malleic Acid in a stoppered container.
- Solid Sodium Hydrogen Carbonate in a stoppered container.

Access to:-

- 2M Sodium Hydroxide supplied with a dropper.
- 2M Ammonium Hydroxide with a dropper.
- 0.05M Potassium Iodide solution with a dropper.
- Acidified Potassium Dichromate (VI) with a dropper.
- Acidified potassium manganate (VII) supplied with a dropper.
- Phenolphthalein indicator supplied with a dropper.
- Bunsen burner.

Notes:

1. Solution Q is prepared by dissolving 4g of sodium hydroxide in distilled water and diluting to one litre.
2. Solution L is prepared by dissolving 3.16g of potassium manganate (VII) in 200cm³ of 2MH₂SO₄ and adding water to make up to one litre.
3. Acidified potassium dichromate (VI) is prepared by dissolving 29.4g in 400cm³ of 2MH₂SO₄ and adding water to make up to one litre.