**KENYAPLEX EXAMS**

**END TERM 3 2019**

**FORM 3**

**CHEMISTRY PAPER 3**

**CONFIDENTIAL INSTRUCTIONS.**

In addition to the apparatus and fittings found in a chemistry laboratory each candidate will require the following:

- Solid P (1.8g of oxalic acid) in a stoppered container.

- Solution Q – 100cm3

- Pipette and pipette filler.

- 100 cm3 measuring cylinder.

- 250cm3 beakers (two)

- 250cm3 volumetric flasks.

- 1 label

- Burette

- One boiling tube

- Five dry clean test tubes in a rack

- Clamp stand

- Solution L – 50cm3

- Stop watch

- 10cm3 measuring cylinder (2)

- Tripod stand and wire gauze

- Test tube holder

- Funnel

- 250 cm3 conical flasks (2)

- Test tube holders

- Distilled water

- White tile

- Solid D – 2g Lead II Nitrate in a stoppered container.

- Solid M – 2g Sodium Chloride in a stoppered container.

- Solid T – 2g Malleic Acid in a stoppered container.

- Solid Sodium Hydrogen Carbonate in a stoppered container.

**Access to:-**

- 2M Sodium Hydroxide supplied with a dropper.

- 2M Ammonium Hydroxide with a dropper.

- 0.05M Potassium Iodide solution with a dropper.

- Acidified Potassium Dichromate (VI) with a dropper.

- Acidified potassium manganate (VII) supplied with a dropper.

- Phenolphthalein indicator supplied with a dropper.

- Bunsen burner.

**Notes:**

1. Solution Q is prepared by dissolving 4g of sodium hydroxide in distilled water and diluting to one litre.

2. Solution L is prepared by dissolving 3.16g of potassium manganate (VII) in 200cm3 of 2MH2S04 and adding water to make up to one litre.

3. Acidified potassium dichromate (VI) is prepared by dissolving 29.4g in 400cm3 of 2MH2S04 and adding water to make up to one litre.