

# **KENYAPLEX EXAMINATION -2019**

## **END OF TERM 1 EXAM**

### **CHEMISTRY PAPER 3**

#### **FORM 4**

#### **CONFIDENTIAL**

**In addition to the apparatus found in the laboratory each candidate will require the following;**

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- About 1g of solid G
- 6 clean test-tubes
- Universal indicator solution and a pH chart
- Ethanol supplied with a dropper
- Clean dry metallic spatula
- 1 boiling tube
- Distilled water
- Solution J, about 130cm<sup>3</sup>
- Solution Q, about 160cm<sup>3</sup>
- Solution R, about 30cm<sup>3</sup>
- Screened methyl orange indicator
- Methyl orange indicator
- 100ml measuring cylinder
- Filter paper
- Means of labeling
- Solid P
- Thermometer
- 100ml beaker

#### ***Access to the following;***

- ❖ Ethanol supplied with a dropper
- ❖ Concentrated sulphuric (VI) acid supplied with a dropper bottle
- ❖ Acidified Potassium dichromate (VI) solution
- ❖ Acidified Potassium Manganate (VII) solution.
- ❖ 2M Ba(NO<sub>3</sub>)<sub>2</sub> solution.
- ❖ 2M NaOH solution.
- ❖ 2M HCl acid.
- ❖ Source of heat.

#### ***Preparation***

Solution J is 0.12M HCL, prepared by adding about 800cm<sup>3</sup> of distilled water to 4.05cm<sup>3</sup> of concentrated HCL of density 108g/cm<sup>3</sup> and making it to one litre of solution

Solution Q is prepared by dissolving 5.3g of anhydrous sodium carbonate in enough distilled water and making up to one litre of solution.

Solution R is prepared by dissolving 15.75g of hydrated barium hydroxide in enough distilled water and top up to one litre of solution.

Solid P is 2.0g of oxalic acid weighed accurately and supplied in a stoppered container

Solid G is sodium sulphite