***KENYAPLEX EXAMINATION -2019***

***END OF TERM 1 EXAM***

***CHEMISTRY PAPER 3***

***FORM 4***

***CONFIDENTIAL***

**In addition to the apparatus found in the laboratory each candidate will require the following;**

* About 1g of solid G
* 6 clean test-tubes
* Universal indicator solution and a pH chart
* Ethanol supplied with a dropper
* Clean dry metallic spatula
* 1 boiling tube
* Distilled water
* Solution J, about 130cm3
* Solution Q, about 160cm3
* Solution R, about 30cm3
* Screened methyl orange indicator
* Methyl orange indicator
* 100ml measuring cylinder
* Filter paper
* Means of labeling
* Solid P
* Thermometer
* 100ml beaker

***Access to the following;***

* Ethanol supplied with a dropper
* Concentrated sulphuric (VI) acid supplied with a dropper bottle
* Acidified Potassium dichromate (VI) solution
* Acidified Potassium Manganate (VII) solution.
* 2M Ba(NO3)2 solution.
* 2M NaOH solution.
* 2M HCl acid.
* Source of heat.

***Preparation***

Solution J is 0.12M HCL, prepared by adding about 800cm3 of distilled water to 4.05cm3of concentrated HCL of density 108g/cm3 and making it to one litre of solution

Solution Q is prepared by dissolving 5.3g of anhydrous sodium carbonate in enough distilled water and making up to one litre of solution.

Solution R is prepared by dissolving 15.75g of hydrated barium hydroxide in enough distilled water and top up to one litre of solution.

Solid P is 2.0g of oxalic acid weighed accurately and supplied in a stoppered container

Solid G is sodium sulphite